

★ KEY ★

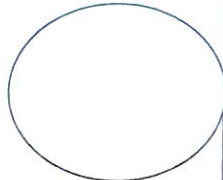
Physical Science – Milestone Review Part 1 -Mixed Review Word Problems

Directions: Read each word problem.

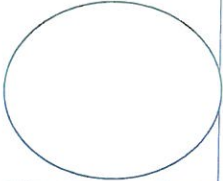
- **Step 1:** Write down the correct number of the formula that you would use to solve the problem.
- **Step 2:** Solve the selected problems on the back of this paper. Show your work. Circle your answer.
- **Use circle diagrams to help you!!!!**

Formula	Word Problem
$F = ma$	1. Calculate the <u>force</u> on an object that has a mass of <u>12 kg</u> and an acceleration of <u>4 m/s²</u> .
$PE = mgh$	2. A bird carries a <u>25 g</u> oyster to a height of <u>11 m</u> . What is the gravitational <u>potential energy</u> of the oyster?
$D = m/v$	3. A loaf of bread has a volume of <u>2270 cm³</u> and a mass of <u>454 g</u> . What is the <u>density</u> of the bread?
$a = \frac{v_f - v_i}{t}$	4. During a race, a sprinter <u>increases speed</u> from <u>5.0 m/s</u> to <u>7.5 m/s</u> over a period of <u>1.25 s</u> . What is the sprinter's <u>average acceleration</u> during this period?
$Q = mc\Delta T$	5. Calculate the <u>specific heat capacity</u> of an unknown substance given that <u>204.75 J</u> of energy raises the temperature of <u>18g</u> of copper from <u>32°</u> to <u>48°C</u> .
$v = d/t$	6. A sprinter runs the <u>100-meter</u> race in <u>10 seconds</u> . What is the sprinter's <u>speed</u> ?
$W = Fd$	7. Your family is moving to a new apartment. While lifting a box <u>1.5m</u> straight up to put it on a truck, you exert an upward force of <u>200 N</u> . How much <u>work</u> did you do?
$v = d/t$	8. A car travels <u>500 miles</u> east in <u>10 hours</u> . What is the <u>speed</u> of the car?
$F = ma$	9. A <u>3,000-N</u> force acts on a <u>200-kg</u> object. The <u>acceleration</u> of the object is ____.
$F = ma$	10. Given a force of <u>88 N</u> and an acceleration of <u>4 m/s²</u> , what is the <u>mass</u> ?
$D = m/v$	11. A block of wood has a density of <u>0.6 g/cm³</u> and a volume of <u>1.2 cm³</u> . What is the <u>mass</u> of the block of wood? Be careful!
$MA = \frac{d_e}{d_r}$	12. Calculate the <u>mechanical advantage</u> of a ramp that is <u>6.0 m</u> long and <u>1.5 m</u> high.
$v = d/t$	13. Sally drove at a speed of <u>50 km /hr.</u> south for <u>2 hours</u> . <u>How far</u> did she travel?
$w = mg$	14. What is the <u>weight</u> of an object that has a mass of <u>25 kg</u> ?
$W = Fd$	15. A young man exerted a force of <u>9,000 N</u> on a stalled car but was <u>unable to move it</u> . How much <u>work</u> was done?
$F = ma$	16. What is the <u>acceleration</u> of a <u>10 kg</u> mass pushed by a <u>5 N</u> force?
$W = Fd$	17. A crane uses an average force of <u>5200 N</u> to lift a girder <u>25 m</u> . How much <u>work</u> does the crane do on the girder?
$a = \frac{v_f - v_i}{t}$	18. A person is traveling at <u>20 m/s</u> in a car when the car hits a tree. The person comes to a <u>complete stop</u> in <u>0.4 seconds</u> . Calculate the <u>acceleration</u> .
$W = Fd$	19. A crane uses an average force of <u>5200 N</u> to lift a girder <u>25 m</u> , but the girder does NOT move. How much <u>work</u> does the crane do on the girder?
$KE = \frac{1}{2}mv^2$	20. A cheetah can run briefly with a speed of <u>31 m/s</u> . Suppose a cheetah with a mass of <u>47 kg</u> runs at this speed. What is the cheetah's <u>kinetic energy</u> ?
$w = mg$	21. What is the <u>mass</u> of an object that <u>weighs 58 N</u> ?
$v = d/t$	22. Find the <u>speed</u> of a long-distance runner who runs <u>30 miles</u> in <u>6 hours</u> .
$Q = mc\Delta T$	23. Calculate the <u>amount of heat</u> needed to increase the temperature of <u>155g</u> of water from <u>23°C</u> to <u>44°C</u> . (The specific heat of water is <u>4.18 J/g°C</u>)
$a = \frac{v_f - v_i}{t}$	24. A plane stops after traveling <u>125 mph</u> in <u>0.5 hours</u> . What is the plane's <u>acceleration</u> ?
$Q = mc\Delta T$	25. <u>241 J</u> of energy is required to raise the temperature of an unknown substance from <u>23°</u> to <u>42°C</u> . Calculate the <u>mass</u> of the substance. (Specific Heat Capacity of the substance is <u>0.90 J/g°C</u>).

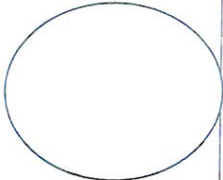
1. Calculate the force on an object that has a mass of 12 kg and an acceleration of 4 m/s^2 .

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

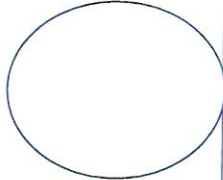
2. A bird carries a 25 g oyster to a height of 11 m. What is the gravitational potential energy of the oyster?

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

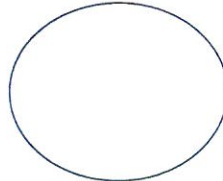
3. During a race, a sprinter increases speed from 5.0 m/s to 7.5 m/s over a period of 1.25 s. What is the sprinter's average acceleration during this period?

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

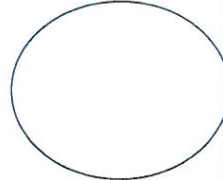
4. Your family is moving to a new apartment. While lifting a box 1.5m straight up to put it on a truck, you exert an upward force of 200 N. How much work did you do?

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

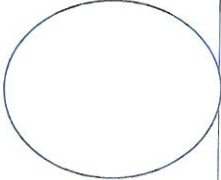
5. A block of wood has a density of 0.6 g/cm^3 and a volume of 1.2 cm^3 . What is the mass of the block of wood? *Be careful!*

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

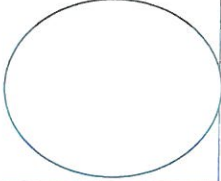
6. Calculate the mechanical advantage of a ramp that is 6.0 m long and 1.5 m high.

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

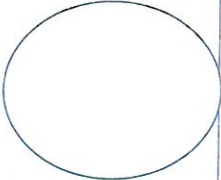
7. What is the weight of an object that has a mass of 25 kg?

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

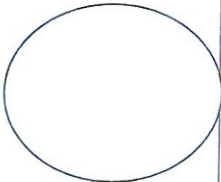
8. A young man exerted a force of 9,000 N on a stalled car but was unable to move it. How much work was done?

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

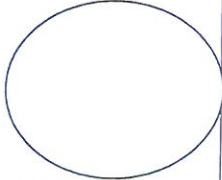
9. A cheetah can run briefly with a speed of 31 m/s. Suppose a cheetah with a mass of 47 kg runs at this speed. What is the cheetah's kinetic energy?

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

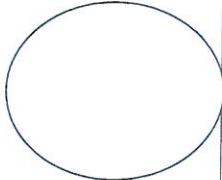
10. Find the speed of a long-distance runner who runs 30 miles in 6 hours.

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

11. Calculate the amount of heat needed to increase the temperature of 155g of water from 23°C to 44°C. (The specific heat of water is 4.18 J/g°C)

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

12. 241 J of energy is required to raise the temperature of an unknown substance from 23° to 42°C. Calculate the mass of the substance. (Specific Heat Capacity of the substance is 0.90 J/g°C).

<i>Given:</i>	<i>Formula:</i>	
	<i>Work:</i>	

Name: ★ KEY ★

Date: _____

Period: _____

Physics Vocabulary Milestone review part 2

What kind of thermal transfer? 1. Conduction, 2. Convection, 3. Radiation		Friction: Match term to description	
<u>3</u> From the sun	<u>3</u> from electromagnetic radiation	1. Rolling Friction	a. Resistance of a fluid on an object
<u>1</u> Touching something		2. Air Friction	b. Resistance of air pushing against object
<u>2</u> In a pot of water	<u>2</u> in moving fluids	3. Fluid Friction	c. Resistance of two object pushing against each other
<u>2</u> Liquids and gases become less dense when hot	<u>1</u> putting your hand on a hot car	4. Sliding Friction	d. Any force that resist motion
	<u>2</u> fan cooling you down.	5. Friction	e. Resistance of a wheel.
Which of Newton's Three Laws Applies?		_____ a heavier animal has to use more muscle to speed up	
<u>3</u> a paddle-wheel boat pushes on the water and the water pushed back to move the boat		_____ you push on the wall and you don't move	
<u>2</u> a tractor trailer truck takes longer to accelerate		_____ Fighter pilot feels massive amount of force when their plane turns quickly	
<u>2</u> a rolling ball hit your leg hard to stop		<u>1</u> a ball won't move until it is kicked	

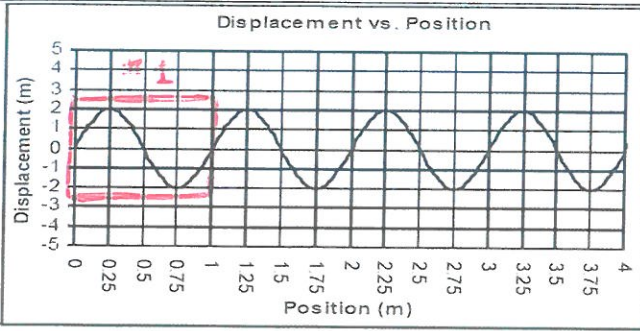
Newton's Laws and Simple Machines Matching

1. Inertia	a. An action that causes motion	1. Weight	a. When all forces on an object are balanced
2. Friction	b. Force pulling all objects towards each other.	2. Equilibrium	b. Force of gravity on an object
3. Gravity	c. Any force that resist motion	3. Mass	c. Acceleration of gravity
4. Net Force	d. Total of all forces on an object	4. Heat	d. Product of friction
5. Force	e. Ability of an object to resist change of motion.	5. g	e. Measure of the matter in an object

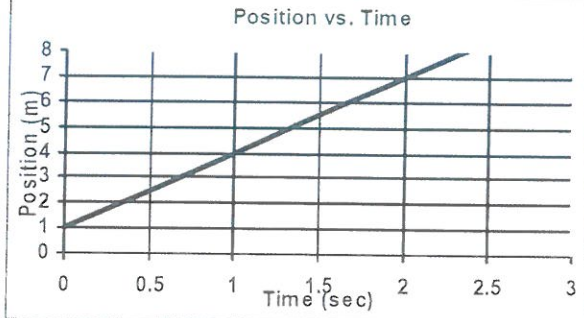
Work and Energy Vocabulary Matching

1. Energy	a. Uses energy and can create energy	1. Thermal	a. Energy of the atom being split or fused
2. Power	b. Energy of motion, dependent on mass and velocity	2. Nuclear	b. Energy cannot be destroyed or created, just transformed
3. Work	c. Energy of position, dependent on height, mass and gravity	3. Mechanical	c. Energy of moving electrons
4. Kinetic Energy	d. The rate of doing work, how fast you do work	4. Law of Conservation of energy	d. Heat energy
5. Potential Energy	e. Has the ability to create forces, stored work.	5. Chemical	e. Energy store in objects and can do work
		6. Electrical	f. Energy of molecular bonds

Interpreting Graphs and Pictures

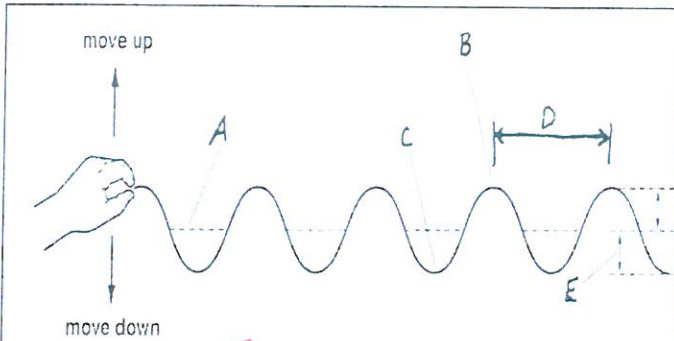


1. Mark one cycle on the wave above
2. Starting at 0.75m, where does the 2nd cycle end? 2.75
3. How many cycles are in the graph? 4
4. Calculate the length of one wave? 1m
5. Calculate Amplitude of wave? 2m

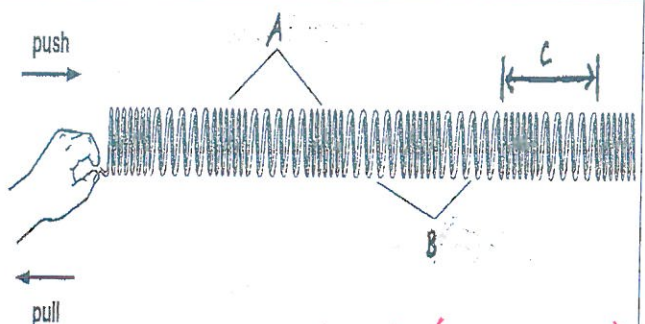


1. What is the position of the object at 2s? 7m
2. When did the object reach 4m? 1 sec
3. Find the speed of the object (show work)
 $s = d/t = 6m / 2sec = 3m/s$

Waves: identify type of wave and label the parts of each wave.



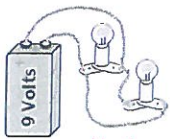
- Wave type: Transverse
- A midline B crest
 C trough D wavelength
 E amplitude



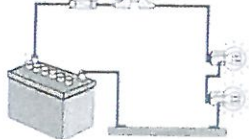
- Wave Type: Longitudinal (compressional)
- A compression
 B rarefaction
 C amplitude

Electricity and Machines

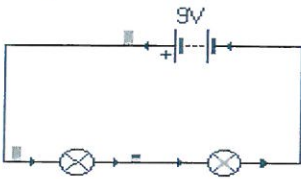
Identify circuits as Series or Parallel



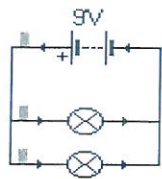
a. parallel



b. series

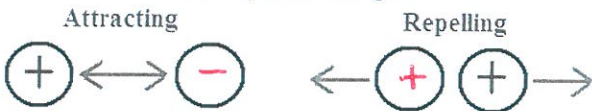


c. series



d. parallel

Complete Charges



Name and give an example of each simple machine

1. lever
Example: _____
2. pulley
Example: _____
3. wheel and axle
Example: _____
4. inclined plane
Example: _____
5. wedge
Example: _____
6. screw
Example: _____

Name: ★ KEY ★

Date: _____

Period: _____

Milestone Review part 3 - Chemistry - Chemical Reactions and Properties of Matter

1. Find the density of the following objects showing your work

Object	Mass	Volume	Density = mass/volume
Large cylinder	28 g	5 ml	$28g/5ml = 5.6g/ml$
Small cylinder	15 g	3 ml	$15g/3ml = 5g/ml$
Small cube	27 g	6 cm ³	$27g/6cm^3 = 4.5g/cm^3$

2. Which two samples of the three types of items above do you think is made of the same material? Why?

3. What is the difference between an ionic compound and a covalent compound?

ionic - donate electrons, covalent - share electrons

4. What is a valence electron?

electrons in outermost shell/energy level

5. What is an oxidation number?

the number of electrons lost (positive number) or gained (negative number)

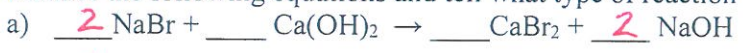
6. What is the difference between a cation and an anion?

cation - positive charge, anion - negative charge

7. Fill in the missing information:

Ionic/Covalent	1 st element w/ charge if ionic	2 nd element w/ charge if ionic	Chemical formula	Chemical Name
<i>I</i>	Na ⁺¹	Cl ⁻¹	NaCl	Sodium Chloride
<i>I</i>	K ⁺¹	S ⁻²	<i>K₂S</i>	<i>Potassium sulfide</i>
<i>I</i>	Ca ⁺²	Cl ⁻¹	<i>CaCl₂</i>	<i>Calcium chloride</i>
<i>C</i>	C	O	<i>CO₂</i>	Carbon dioxide
<i>C</i>	N	O	N ₂ O ₅	<i>Dinitrogen pentoxide</i>
<i>I</i>	Mg ⁺²	O ⁻²	<i>MgO</i>	Magnesium oxide
<i>C</i>	S	O	<i>SO₃</i>	Sulfur trioxide
<i>I</i>	Mg ⁺²	P ⁻³	<i>Mg₃P₂</i>	<i>Magnesium phosphide</i>
<i>I</i>	Al ⁺³	O ⁻²	<i>Al₂O₃</i>	<i>Aluminum oxide</i>
<i>C</i>	<i>O</i>	<i>F</i>	<i>OF₃</i>	Oxygen trifluoride
<i>C</i>	<i>C</i>	<i>Cl</i>	CCl ₄	<i>Carbon tetrachloride</i>
<i>I</i>	Al ⁺³	Cl ⁻¹	<i>AlCl₃</i>	<i>Aluminum chloride</i>
<i>I</i>	Ca ⁺²	O ⁻²	<i>CaO</i>	<i>Calcium oxide</i>
<i>C</i>	<i>P</i>	<i>O</i>	P ₂ O ₅	<i>Diphosphorus pentoxide</i>
<i>I</i>	Na ⁺¹	S ⁻²	<i>Na₂S</i>	<i>Sodium sulfide</i>

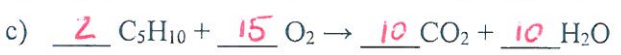
8. Balance the following equations and tell what type of reaction it is



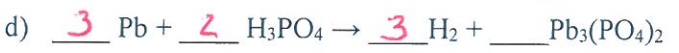
reaction type double displacement



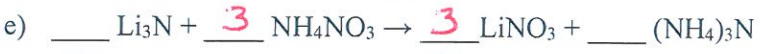
reaction type synthesis



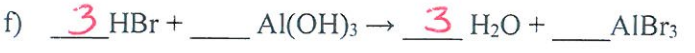
reaction type combustion



reaction type single displacement



reaction type double displacement



reaction type double displacement

Milestone Review part 4 Chemistry—Atomic and Nuclear Theory and the Periodic Table

1. List the three subatomic particles; tell where they are found and what charge they have.

Particle	location	charge
Proton	Nucleus	Positive (+)
Neutron	Nucleus	Neutral (0)
Electron	Electron Cloud	Negative (-)

2. Label the parts of the beryllium (Be) atom to the right ----->

3. What is an isotope? *element that has a different mass than the stable*

4. How are the elements arranged in the periodic table? *increasing atomic #*

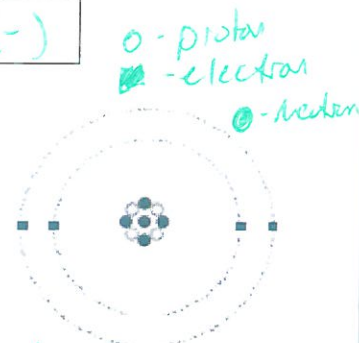
5. Valence electrons determine how an atom will react.

6. What does the atomic number tell you? *# of protons*

7. How do you find the number of neutrons an atom of an element would have? *Mass # - atomic #*

8. What are valence electrons and how do you find out how many valence electrons an element has?

*Electrons in the outermost energy shell. Their group # tells you
(group 2 has 2; group 13 has 3)*



9. Give the number of protons, neutrons and electrons in the following isotopes:

	# of protons	# of neutrons	# of electrons
a. Mg-24	12	12	12
b. Mg-26	12	14	12
c. N-15	7	8	7
d. O-18	8	10	8
e. Si-30	14	16	14
f. S-34	16	18	16

10. Complete the table. There is enough information given for each element to determine all missing numbers

Symbol	Atomic Number	Mass Number	Number of Protons	Number of Electrons	Number of Neutrons
²³ Na	11	23	11	11	12
K	19	40	19	19	21
F	9	19	9	9	10
Ca	20	41	20	18	21
¹³¹ I	53	131	53	53	78
Ag	47	109	47	46	62
H	1	2	1	1	1
³⁶ S	16	36	16	16	20

11. What is radioactivity? spontaneous decay of an unstable isotope
12. What is half-life? the time required for 1/2 of a radioactive substance to decay
13. If we start with 400 atoms of a radioactive substance, how many would remain after one half-life? 200
 after two half-lives? 100 after three half-lives? 50 after four half-lives 25?
14. A paleontologist discovered fossil remains of ancestral mammal, and in order to have basis of comparison to other ancestral mammals, he needed the age of the fossil. The carbon-14 analysis indicates that only half of the original amount is present. How old is this fossil? 5,730 years
15. The half-life of hydrogen-3 is 12.3 years. Given 100 g of hydrogen-3, how many grams will be left after 5 half-lives?
3.125 g
16. A patient is administered 20 mg of iodine-131. How much of this isotope will remain in the body after 40 days if the half-life for iodine-131 is 8 days?
0.625 mg
17. The mass of cobalt-60 in a sample is found to have decreased from 0.8 g to 0.2 g in a period of 10.5 years. From this info, calculate the half-life of cobalt-60
5.25 yrs
18. The three most common states of matter are solid, liquid & gas.
19. The kinetic theory states that the higher the temperature, the (faster / slower) the particles that make up a substance move.
20. As a sample of matter is cooled, its particles move more (slowly / quickly).
21. The particles that make up a solid move slower than do the particles that makes up a gas.
22. Matter that has a definite volume but no definite shape is a liquid.
23. Matter that has a definite volume and a definite shape is a solid.
24. If you move a substance from one container to another and its volume changes, the substance is a gas.



Substance A



Substance B



Substance C

25. In the above picture, which substance is a liquid? B
26. In the above picture, which substance is a solid? A
27. In the above picture, which substance is a gas? C
28. In the above picture, which substance are the forces of attraction among the particles so weak that they can be ignored under ordinary conditions? C

Melting and Boiling Points of Some Substances		
Substance	Melting Point	Boiling Point
Hydrogen	-259.3°C	-252.9°C
Nitrogen	-210.0°C	-195.8°C
Acetic Acid	16.6°C	117.9°C
Gold	1064.2°C	2856°C

29. Based on the information in the table above, the melting point of acetic acid is 16.6°C.
30. Based on the information in the table above, the freezing point of nitrogen is -210°C.
31. Based on the information in the table above, which substances would be a gas at 0°C? H & N

32. A solution is a Chemical mixture of two or more components.
33. The Solvent is the component in the greatest amount.
34. The solute is the component in the least amount
35. In a mixture, the solute dissolves in the solvent.
36. If you dissolve sugar in water, which is the solvent and which is the solute? Water = Solvent Sugar = solute
37. Soft drinks consist of a mixture of water, sugar, and flavoring, with carbon dioxide gas bubbled through it. Which of these ingredients would be considered the solvent? Water
38. Dry air is primarily made up of nitrogen (78.09%) and oxygen (20.95%). Which of these is the solvent and which is the solute? Nitrogen = Solvent Oxygen = solute

Classification of Matter—Know definition of Matter; Be able to give examples of and tell difference between: Mixtures and substances (non-mixtures); Homogenous and heterogeneous mixtures; Elements vs. Compounds.

Label as Mixture or Pure Substance	1. Pure substance <u>D</u>	a. Made up of two types of matter that can be physically separated.
Salt Water <u>M</u> Chicken Soup <u>M</u>	2. Mixture	b. Two samples might not be the same.
Water <u>PS</u> Salt <u>PS</u>	3. Heterogeneous Mixture <u>a</u>	c. Two samples will have the same makeup.
Silver <u>PS</u> Chex mix <u>M</u>	4. Matter <u>g</u>	d. Has only one kind of atom in the sample.
Label as Homogenous or Heterogeneous	5. Element <u>D</u>	e. Contains two kinds of atoms that <i>cannot</i> be physically separated.
Sugar Water <u>M</u> Vegetable Soup <u>E</u>	6. Homogeneous Mixture <u>E</u>	f. Cannot be separated by physical means.
Chex Mix <u>E</u> Jello w/ fruit <u>E</u>	7. Compound <u>E</u>	g. A classification of anything that has mass and takes up space.
Milk <u>M</u> Plain Jello <u>M</u>		

Write the metric prefixes in order from largest to smallest:

Kilo hecto deka deci centi mili

<p>Convert the Following</p> <p>3.2 kilometers = <u>3200</u> meter</p> <p>0.23 centimeters = <u>2.3</u> millimeters</p> <p>0.12 liter = <u>120</u> milliliters</p> <p>2500 millileters = <u>2.5</u> liters</p> <p>4500 grams = <u>4.5</u> kilograms</p> <p>9 kilograms = <u>9000</u> grams</p> <p>54 megaliters = <u>540000000</u> centiliters</p>	<p>Identify Physical or Chemical Change</p> <p>Sugar dissolved in water <u>Phys</u></p> <p>Wood burning <u>Chem</u></p> <p>Digestion <u>Chem</u></p> <p>Water Boiling <u>Phys</u></p> <p>Two liquids bubble when mixed <u>Chem</u></p> <p>Cooking food <u>Chem</u></p> <p>Melting butter <u>Phys</u></p>												
<p>Circle the acids and underline the bases</p> <p><u>H₂(PO₄)</u> <u>HF</u> <u>Ca(OH)₂</u></p> <p><u>Ca(OH)₂</u> <u>NaOH</u> <u>HNO₃</u></p>	<p>Circle the indicators of <u>physical changes</u></p> <p>Underline the indicators of a <u>chemical change</u></p> <table border="1"> <tr> <td><u>Melts</u></td> <td>Produces gas.</td> <td>Changes in color</td> </tr> <tr> <td><u>Changes Smell</u></td> <td><u>Ripped</u></td> <td><u>Cutting</u></td> </tr> <tr> <td><u>Boils</u></td> <td>Turns cloudy</td> <td>Changes taste.</td> </tr> <tr> <td><u>Breaks</u></td> <td><u>Dissolves</u></td> <td><u>freezing</u></td> </tr> </table>	<u>Melts</u>	Produces gas.	Changes in color	<u>Changes Smell</u>	<u>Ripped</u>	<u>Cutting</u>	<u>Boils</u>	Turns cloudy	Changes taste.	<u>Breaks</u>	<u>Dissolves</u>	<u>freezing</u>
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1. Proton C	a. Particles with no charge that exists in the nucleus of most atoms.	1. Atomic Number B	a. Total number of protons and neutrons in the nucleus of an atom.
2. Neutron A	b. Center of the atom, contains most of the atom's mass.	2. Molecule D	b. Number of protons in an atom; also the way the elements are numbered.
3. Electron F	c. Positively charged particle in the nucleus of the atom.	3. Compound E	c. An atom with a different number of neutrons
4. Nucleus B	d. The smallest part of an element or molecule.	4. Mass Number A	d. Two or more elements combined.
5. Atom D	e. Negative particles in the nucleus of the atom.	5. Isotope C	e. Two or more atoms that are combined
	f. Negatively charged particle that exists in the space around the nucleus.		f. Number of electrons in an atom.

Solution (So); Colloid (C); Suspension (Sp)	The temperature at which a solid turns to liquid is called: Melting
Milk in water _____ Doesn't settle; scatters light _____	The temperature at which a liquid turns to a gas is called: boiling / evaporation
Vinegar in water So Doesn't scatter light or settle _____	The temperature at which a gas turns to liquid: condensation
Sand in water Sp Settles and scatters light _____	The temperature at which a liquid turns to a solid: freezing
Oil and water _____ Fog _____ Milk _____	When a solid turns straight to a gas is called: sublimation
	At what temperature does water melt? 0°C
	At what temperature does water boil? 100°C

Endothermic or Exothermic	Give the group/family name for the following groups on periodic table
1. If it gets cold Endo	Group 1: Alkali line metals
2. If it gets hot Exo	Group 2: Alkali earth metals
3. Condensation: Endo	Group 13: Baron group
4. Vaporization: Exo	Group 14: Carbon group
5. If it absorbs heat Endo	Group 15: Nitrogen group
6. If it releases heat Exo	Group 16: Oxygen group
7. Melting Endo	Group 17: Halogens
8. Freezing Exo	Group 18: Noble gases

Reading Chemical Equations

$\text{Li}_2\text{O} + \text{MgCl}_2 \rightarrow 2 \text{LiCl} + \text{MgO}$	$2 \text{C}_3\text{H}_7\text{OH} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 8 \text{H}_2\text{O}$
Write the second reactant: Magnesium chloride	Write the first product: Carbon Dioxide
Write the first product: Lithium chloride	How many carbons on the reactant side? 6

How many Lithium atoms are on the product side?

2

What is coefficient for lithium chloride? 2

Type of reaction: Double Displacement

How many hydrogens on the product side?

16

What is the coefficient for carbon dioxide?

6

Type of reaction: Combustion